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Graham's Law. • speed of diffusion of 2 different gases is inversely proportionality to the square root of their densities. Raoult's Law. • vapor pressure lowering due to effects of electrolytes or solutes in solution. • the partial vapor pressure of the solvent of an ideal solution is equal to the product of the mole fraction of the solvent and its vapor pressure in pure state.

Physical Pharmacy Notes | Intermolecular Force | Gases

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Dr Valizadeh (Teaching Materials) - Pharmacy faculty

Objectives of the lecture After completion of thislecture, the student should be able to: 1. Understand the various types of pharmaceutical solutions. 2. Define solubility, saturated & unsaturated solutions and polar & non polar solvents. 3. Describe complete & partial miscibility. 4.

Physical Pharmacy- lecture 1

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Martin's Physical Pharmacy Book PDF » StudyFmrd

Physical pharmacy Physical pharmacy. 1. PHYSICAL PHARMACY. 2. Matter: It's what the world is made of. 3. Kinetic Theory of Matter Matter is made up of particles which are in continual random motion. 4. Binding Forces between Molecules The Three States of Matter: In order for molecules to exist in aggregates in gases, liquids and solids Intermolecular forces must exist 4.

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• Physical properties - a characteristic that can be observed or measured without changing the identity or composition of the substance • Physical properties used to describe matter can be classified as: 1) Extensive – depends on the . amount. of matter in the sample - e.g. Mass, volume, length . 2) Intensive – depends on the . type. of

THE PROPERTIES AND STRUCTURE OF MATTER

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