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When matters change state from liquid to gas, they absorb

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energy. I is endothermic reaction. H_1 is positive. In decomposition reactions energy (heat) is absorbed. III is endothermic reaction. H_2 is positive.

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Choose the one alternative that best completes the statement or answers the question. ΔH $4\text{Al (s)} + 3\text{O}_2 \text{ (g)} \rightarrow$

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2Al₂O₃ (s) A) exothermic, released B) exothermic, absorbed
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Name _____ MULTIPLE CHOICE.

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Answer: Change in internal energy (ΔU) for an isolated system is zero for it does not exchange any energy with the surroundings. But entropy tends to increase in case of

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spontaneous reaction. Therefore, $\Delta S > 0$ or positive.
Question 17.

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The pretension is by getting solution and thermochemistry test answer keys as one of the reading material. Kinetic Energy: Energy due to the motion of the object $\frac{1}{2}mv^2$ KE mv
Law of Conservation of Energy: Energy can neither be created nor destroyed,.

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completes the statement or answers the question 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure
Ap Chemistry ...

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Hey there, science buffs! If you specialize in the subject of thermochemistry then these quizzes are for you! Pull on your white coats and safety goggles, we ' re about to have some fun! And for those of you who don ' t know what thermochemistry is but would love to give these quizzes a try: thermochemistry is the study of the energy and heat associated with chemical reactions.

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Solution In order to obtain the smallest amount of heat in which there is a large decrease in the number of moles of gas, you should carry out the reaction under conditions of constant volume. According to Avogadro ' s Law ($V_1/n_1 = V_2/n_2$), volume is directly proportional to the number of gas molecules.

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